

Figure 1 is a schematic diagram of the experimental setup for the study of the effect of the initial concentration of the reactants on the rate of the reaction. The diagram shows a reaction vessel containing a mixture of reactants, with arrows indicating the flow of reactants into the vessel and the flow of products out. The reaction is labeled as $A + B \rightarrow C + D$. The initial concentration of the reactants is denoted as $[A]_0$ and $[B]_0$. The rate of the reaction is denoted as v . The diagram also shows the effect of the initial concentration of the reactants on the rate of the reaction, with a graph of v vs. $[A]_0$ and $[B]_0$. The graph shows that the rate of the reaction increases with the initial concentration of the reactants.

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